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Dalton's atomic theory. theory proposed  
by John Dalton in 1808 which states that

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matter is composed of extremely small particles called atoms; marked the beginning of modern atomic theory.

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Section 4.4 Changes to the Nucleus—Nuclear Reactions In your textbook, read about radioactivity. For each item in Column A, write the letter of the matching item in Column B.

Column A

1. The rays and particles that are emitted by a radioactive material
2. A reaction that involves a change in an atom's nucleus
- 3.



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Glencoe Chemistry Matter and Change - Chapter 4. Dalton's atomic theory. Atom. Cathode Ray. Electron. States that matter is composed of extremely small

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particles ca.... Smallest particle of an  
element that retains all the propertie....  
Radiation that originates from the  
cathode and travels to the....

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Answer Key

SOLUTIONS MANUAL 37. Conservation of Mass Explain how Dalton's atomic theory offered a convincing explanation of the observation that mass is conserved in chemical reactions. Mass is conserved because atoms cannot be created, divided, or destroyed. Chemical reactions

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**The Structure of the Atom**  
**The Structure of the Atom**

Chapter 4 Solutions; Chapter 4 review problems 9/6/13; Chapter 6 - Thermodynamics. Chapter 6 Study Guide; Chapter 6 Outline; Chapter 6 Assignments; Chapter 6 homework problems; Chapter 11 - Intermolecular Attractive Forces and Liquids and Solids.

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Chapter 11 Study Guide; Chapter 11  
Outline notes; Phase Change Problem  
from 9/20/2013; Chapter 12 ...

### **Baylor, Scott / Chemistry: Matter and Change**

section 4.4 Unstable Nuclei and  
Radioactive Decay In your textbook,  
read about radioactivity. For each item

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in Column A, write the letter of the matching item in Column B. Column A

1. The rays and particles that are emitted by a radioactive material
2. A that involves a change in an atom's nucleus
- 3.

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Chemistry: Matter and Change Chapter 8  
44 . Name Date CHAPTER FOR Class  
Section 8.2 continued 9 What is the  
relationship between lattice energy and  
the strength of the attractive force  
holding ions in place? a. The more  
positive the lattice energy is, the greater

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the force. b. The more negative the lattice energy is, the greater the force.

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Chapter 3 Solutions Manual CHAPTER 3  
SOLUTIONS MANUAL 6. From a  
laboratory process designed to separate

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water into hydrogen and oxygen gas, a student collected 10.0 g of hydrogen and 79.4 g of oxygen. How much water was originally involved in the process?  
mass reactants mass products mass reactants mass ...

**Matter—Properties and  
Changes Matter—Properties and**

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**Changes**

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2. chemical change (4/2) 3. conservation  
of mass (3/2) 4. physical change (4/2) 5.

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## Answer Key

density (1/1) 6. physical property (1/1) 7.  
boiling point (1/1) 8. behavior (1/1) 9.  
volume (1/1) 10. melting point (1/1) 11.  
state (2/1) 12. appearance (1/1) Part B.  
Concept Review (page 43) 1. physical  
properties (1/1) 2. chemical change (4/2)  
3. chemical (4/2 ...

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Supplemental Problems Chemistry:

Matter and Change • Chapter 3 3

Matter—Properties and

Changes Properties and Changes 1. An 18-g sample of element A combines completely with a 4-g sample of element B to form the compound AB. What is the mass of the compound formed? 2. A substance breaks down into three



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component elements when it is ...

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